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Note: Not all of the reactions will occur.
For those that do not, write no reaction.

1. A piece of copper is dropped into a
container of water. No reaction
2. Liquid bromine is added to a container of
sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)}$

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$2\text{NaBr}(s) + 12(s)$ 3. An aluminum strip is immersed in a solution of silver nitrate.

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oxidation number (valence or charge) equal to zero. O, Fe, T, He < /1 1 ~ ~ L-td. 8. The following are diatomic or polyatomic elements in nature which must be committed to memory.

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do not break apart and are in net ionic
equation. • Synthesis and
Decomposition Reactions - Pgs. 42-43,
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Handbook SnCO tin(II) carbonate sodium
hydrogen carbonate NaHCO 1. 2. 3. 4. 5.
6. 8. 9. 10. 11. vanadium(V) oxide H₂O
dihydrogen monoxide (NH) CO

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ammonium oxalate 4 224 polonium(VI)
thiocyanate $\text{Po}(\text{SCN})_6$ tetraphosphorus
decaoxide P_4O_{10} 12. 13. 15. 16. 17.
18. 19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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in part (ii) answer the question about the
reaction.

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number of the underlined element:
KMnO₄ Since K is an alkali metal, its

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charge must be 1+ Oxygen is 2— but there are four of them, therefore, 4 times 2— equals 8— If 1+ and 8— are added Ultimate Chemical Equation Answers

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 KMnO_4 Since K is an alkali metal, its charge must be $1+$ Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$ If $1+$ and $8-$ are

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