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Answers

# Modern Chemistry Chapter 3 Section Answers

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Law 3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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## *1 Matter and Change - HUBBARD'S HONORS CHEMISTRY CLASS*

Chapter 6: Chemical Bonding, Modern Chemistry. This is a working presentation of the notes for this chapter. ... Chemical bonding chapter 6  
1. Chapter 6 - Modern CHEMICAL BONDING Chemistry S.Martinez ... IONIC BONDS AND Chapter 6 Section 3 COMPOUNDS 11/19/2010 S.Martinez 39  
40. IONIC BONDING & IONIC COMPOUNDS Most of the rocks & minerals in the ...

## *6 Chemical Bonding*

CHAPTER 5 REVIEW The Periodic Law  
SECTION 1 SHORT ANSWER Answer the following questions in the space provided.  
1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.  
2. d Mendeleev noticed

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that certain similarities in the ...

*Chemical bonding chapter 6 - SlideShare*  
CHAPTER 10 REVIEW States of Matter  
SECTION 3 SHORT ANSWER Answer the  
following questions in the space  
provided. 1. Match description on the  
right to the correct crystal type on the  
left. ... SECTION 5 MODERN CHEMISTRY  
STATES OF MATTER. STATES OF MATTER  
MODERN CHEMISTRY.

## *Modern Chemistry Chapter 3 Flashcards | Quizlet*

see figure 3-3 page 67 . C: Modern  
Atomic Theory. Some of Dalton's  
postulates have been proven to be  
incorrect. His third postulate indicated  
that atoms cannot be subdivided. His  
second postulate indicates atoms of the  
same element have identical mass.  
Isotopes. His theory has been modified --  
not discarded. Homework: Section 3-1

## *Holt McDougal Modern Chemistry Chapter 3 Section 1 ...*

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*Modern Chemistry (Florida 2015-2016)  
Chapter 3 Section 3 ...*

if two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element is always a ratio of small whole numbers; textbook: if two or more different compounds are composed of the same two elements with a certain mass of the first element is always a ratio of small whole numbers

*Modern Chemistry Chapter 8 Section 3*

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*mc06se cFMSr i-vi*

CHAPTER 1 REVIEW Matter and Change  
MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance.  
homogeneous a. sugar homogeneous d. plastic wrap homogeneous b. iron filings

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heterogeneous e. cement sidewalk  
heterogeneous c. granola bar 2. For  
each type of investigation, select the  
most appropriate ...

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*10 States of Matter - Ms. Agostine's  
Chemistry Page*

MODERN CHEMISTRY The atomic  
number of nickel-60 is 28. How many  
neutrons does this iso- tope have?  
Carbon-14 has 8 neutrons. What is the  
atomic number of carbon-14? 11  
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CHAPTER 6 REVIEW Chemical Bonding

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SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b ...

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## Answers

*Chapter 3 Modern Chemistry - Prepchem*

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Chemistry Chapter 3 Section 1

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and more with flashcards, games, and  
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### *5 The Periodic Law*

3:1 mol N<sub>2</sub>; 2 mol NH<sub>3</sub>:3 mol H<sub>2</sub>; or  
their reciprocals 7. Given the following  
equation:  $4\text{NH}_3(\text{g}) + 6\text{NO}(\text{g}) \rightarrow 5\text{N}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$   
1 mol NO:1 mol H<sub>2</sub>O a. What is  
the mole ratio of NO to H<sub>2</sub>O? 3 mol  
NO:2 mol NH<sub>3</sub> b. What is the mole ratio  
of NO to NH<sub>3</sub>? 0.360 mol c. If 0.240 mol  
of NH<sub>3</sub> react according to the above  
equation, how many moles of NO will ...

### *3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...*

Holt Earth Science Chapter 18, Section  
18.1; HOLT MODERN CHEMISTRY- CHEM  
SAFETY QUIZ DOC; Notes for Starr  
Taaggart's AP Bio Textbook? ...

Chemistry Chapter 11. Chemistry  
Chapter 2. Chemistry Chapter 15.

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