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gas

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continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18.

The decomposition of a substance by an electric current is called . 19. A(n) orders

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the elements by the ease with which they undergo certain chemical reactions.
20.

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Chapter: The Periodic Law Use the
periodic table below to answer the
questions in this Chapter Test. In the

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space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1.

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16. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by great distances. This implies that the

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continued 18. Dissolution processes with
negative enthalpies of solution are
processes. 19. is the solution process

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with water as the solvent. 20. Liquid solutes and solvents that are not soluble in each other are. 21. A solute molecule that is surrounded by solvent molecules is ...

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the rates of effusion of gases at the

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same temperature and pressure are
inversely proportional to the square
roots of their molar masses Rate of
effusion of A / B = $\sqrt{M_b} / \sqrt{M_a}$

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continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.

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Vocabulary-New ... CH 9 Test and CH 11
Test . CH 9 Limiting Reactant-New. CH 9
Limiting Reactant and % Yield
Chapter 11 Physical Chapter 11 Molar
Volume CH 11 Gases & Gas Laws.

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Modern Chemistry 47 Chapter Test

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continued _____ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive

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Answers

forces between ions are greater than the attractive ...

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Name Class Date Chapter Test A,
continued ____19. In the figure on the
previous page, the pH at the

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equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____ 20. In the figure on the previous page, the volume of titration standard

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