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Module A - Chapter 5 Process Oriented
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Configuration POGIL Unit 2 Module B: Chemical Symbols / Introduction to Periodic Table Chapter/Section Topic Ch. 6: 6.1, 6.2, ... Describe how you determined your answer to part a. 3. Examine each diagram in Model 1 and the corresponding manager's code.

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ChemActivity 11 Electron Configurations and the periodic Table 2. Based on the analysis we have used to assign peaks in photoelectron spectra to shells and subshells in atoms, why is the peak at 0.42 W/mole in the K spectrum assigned to the $n = 4$ shell (as opposed to being another subshell of $n = 3$)? Refer to

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4 POGIL™ Activities for High School Chemistry 7. Consider the orbital diagram for oxygen in Model 2. a. How many electrons are present in the orbital diagram? b. Based on its position in the periodic table, explain how you know that your answer to part a is the correct number of electrons for oxygen. 8.

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The p orbital looks like a peanut

(p=peanut). This orbital has 3

orientations and can therefore hold 6

total electrons. The p orbitals and

orientations are shown on the right. The

orientations refer to how many different

ways the orbital can orient itself in

space. This orbital is found in energy

levels 2-7. d Orbital The d orbital looks

like a double peanut (d=double peanut)

and can hold 10 ...

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ChemActivity 10 Electron Configurations 57 ...
Wha:ÿ information is provided by an electron configuration? What is the relationship between the data in Tables

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1 and 2? Is it possible to deduce the electron configuration for an atom spectrum? If so, describe how. ...

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Ground state electron configurations can
be predicted by a strict set of rules
known as the Rules of Aufbau (meaning
filling up). Examine the diagrams in
Model 2 to determine the correct phrase
which

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team providing the electron
configuration gets the point. Halfway
through the game, I give each person an
element, and the game changes so the
opposing team has to provide the
electron configuration. The team at the
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Orbitals. Model 1: Orbitals. The quantum
mechanical model determines the
allowed energies an electron can have
and how likely it is to find the electron in
various locations around the nucleus.

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