

# Chemactivity 6 Atomic Size Answers

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The most commonly used measure of size of an

atom is its bonding atomic radius, also called the covalent radius.<sup>1</sup> The bonding atomic radius of an element is taken as one half the distance between the nuclei when two identical atoms are bonded together. For example, the internuclear distance in  $\text{Cl}_2$  is 198 pm, so the atomic radius is taken to be ...

## Chemactivity 6 Atomic Size Answers

Answers to ChemActivity 1 – The Nuclear Atom-  
Chapter 5 1. 6 2. 6,7,7 3. 6,6,7 4. (a) neutral same e  
and p, ion different e and p (b) assign +1 to each  
proton and -1 to each electron and take the  
difference.

## The Nuclear Atom

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6. a) Define mass number. b) Define atomic



List the following particles in order of decreasing size: Kr, Sr  $2+$ , Rb  $+$ . 15. Compare the sizes of a negative ion and its neutral atom. Illustrate with specific examples. 16. List the following particles in order of decreasing size: K  $+$ , Ar, S  $2-$ , Cl $-$ , Ca  $2+$ . 17. Arrange the following elements in increasing order of their atom's size. Ca, Ba ...

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atomic mass. c) True. Na atoms are more massive, so fewer atoms needed for same total mass. D) False. This would imply that an atom has a mass of 0.045 amu which is not possible as it is less than the mass of a proton.

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Atomic Radii Size: 1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together. 2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali

metals & period #2 on the periodic table to your right. 3.

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