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Chapter 9 Review Stoichiometry Section

CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual

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yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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Modern Chemistry 77 Stoichiometry
CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following

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SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 mol of H

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CHAPTER 9 REVIEW. Stoichiometry.
SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O₂ form if 3.0 mol of KClO₃ are totally consumed? 2. Given the following equation ...

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represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(s) \rightarrow 2\text{KCl}(s) + 3\text{O}_2(g)$... CHAPTER 9 REVIEW ...

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A

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SECTION 2 continued Date Class _____
60.2 g of lithium chloride, LiCl , is dissolved in 100. mL of water in a 250-mL beaker. The oxygen gas produced in part (a) has a density of 1.43 g/L. Calculate the volume of this gas. 76
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notebook, solve the following problems.

SECTION 9.1 THE ARITHMETIC OF

EQUATIONS Use the 3-step problem-solving approach you learned in Chapter 4. 1. An apple pie needs 10 large apples, 2 crusts (top and bottom), and 1 tablespoon of cinnamon.

9 Stoichiometry Practice Problems

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

Download Chapter 9 Review Stoichiometry Section 2 Work

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a

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chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

Stoichiometry Worksheet Answers Chapter 9

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